Review unit 1-Ch1.1, Ch2,Ch3.3,Ch4,Ch5.1, 5.2, Ch 6.1,6.2 Chem 11-Mrs. Sanford

Name: period:

Matter:

Chemistry:

Five areas of Chemistry:

Extensive property

Intensive property

Physical property

States of matter 1- 2-

3-

Physical change: Give 3 ex:1- 2- 3-

Two types of mixtures and a description for each:
1-
2-

Solution: Phase:

Ways to separate matter:

1- 2-
3 4
5 6
7 8

Element: compound:

Chemical change:

A substance differs from a mixture in that the composition:

Chemical property:

Four things that tell us a chemical change has taken place:
1. 2. 3. 4.

Precipitate:

Law of Conservation Of Mass:

Atomic Number- Mass Number-

Number of neutrons=

Isotope: Give an example of two isotopes of Neon →

Scientist	Key points of their discovery
Democritus	
Dalton	
JJ Thomson	
Nagaoka	
Rutherford	
Bohr	
deBroglie	
Shroedinger/Heisenberg	
Chadwick	
Einstein	
Mendeleev	

Diagonal rule: Draw the energy level filling diagram :
AND Give the electron configuration for sulfur and barium

2s
1s

An s orbital can hold a maximum of_____electrons....a p orbital can hold a maximum of_____electrons....d orbital_____electrons and f orbital_____electrons

family	Group or location
Alkali metals	
Alkaline earth metals	
Halogens	
Metals	
Non-metals	
Transition metals	
Lanthanides	
Actinides	
metalloids	

Metric conversion: Draw table of conversion prefixes on the reverse of this page.
Do the following questions p 84 32 all,33b to c, p 87 42 all, 44 all, p 153 #1, p 111 #15, 16, p 112 #17 and p 113 #19